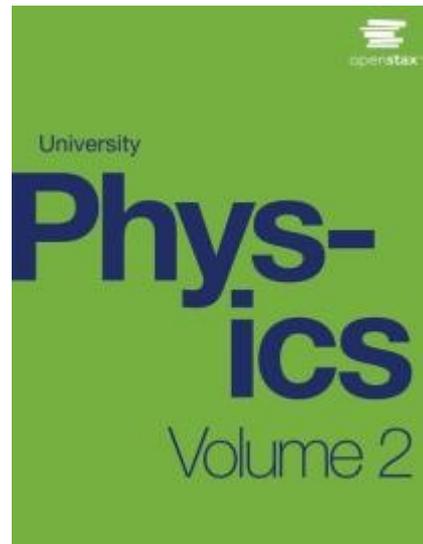


UNIVERSITY PHYSICS

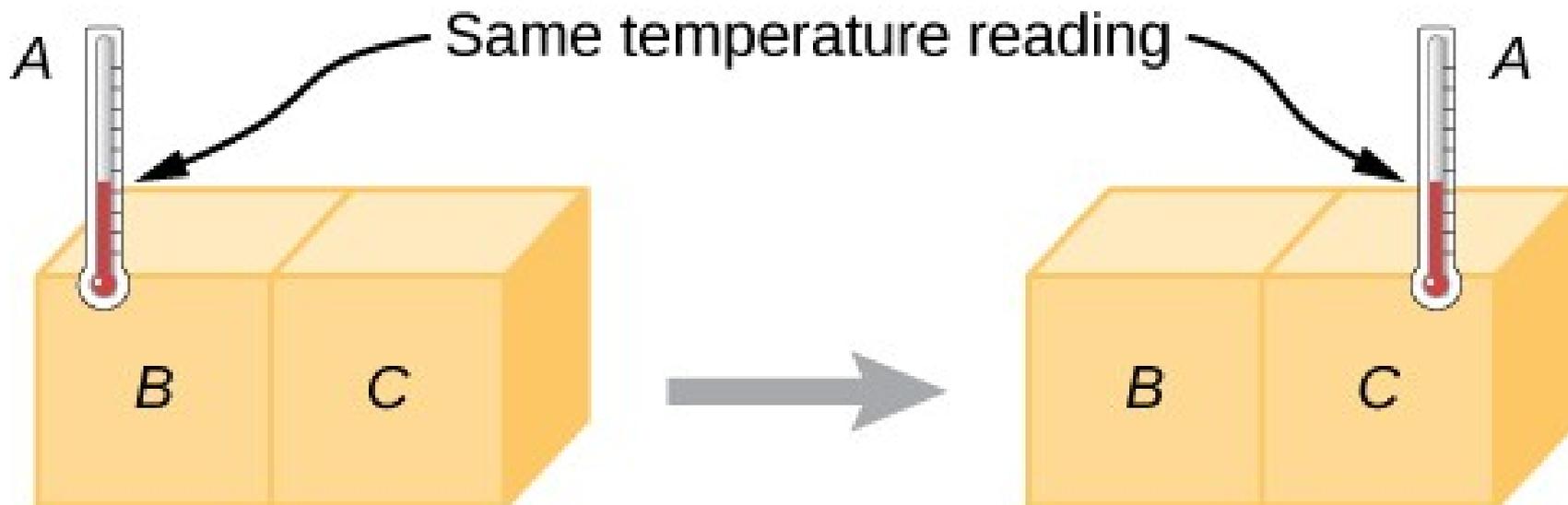
Chapter 1 TEMPERATURE AND HEAT

PowerPoint Image Slideshow



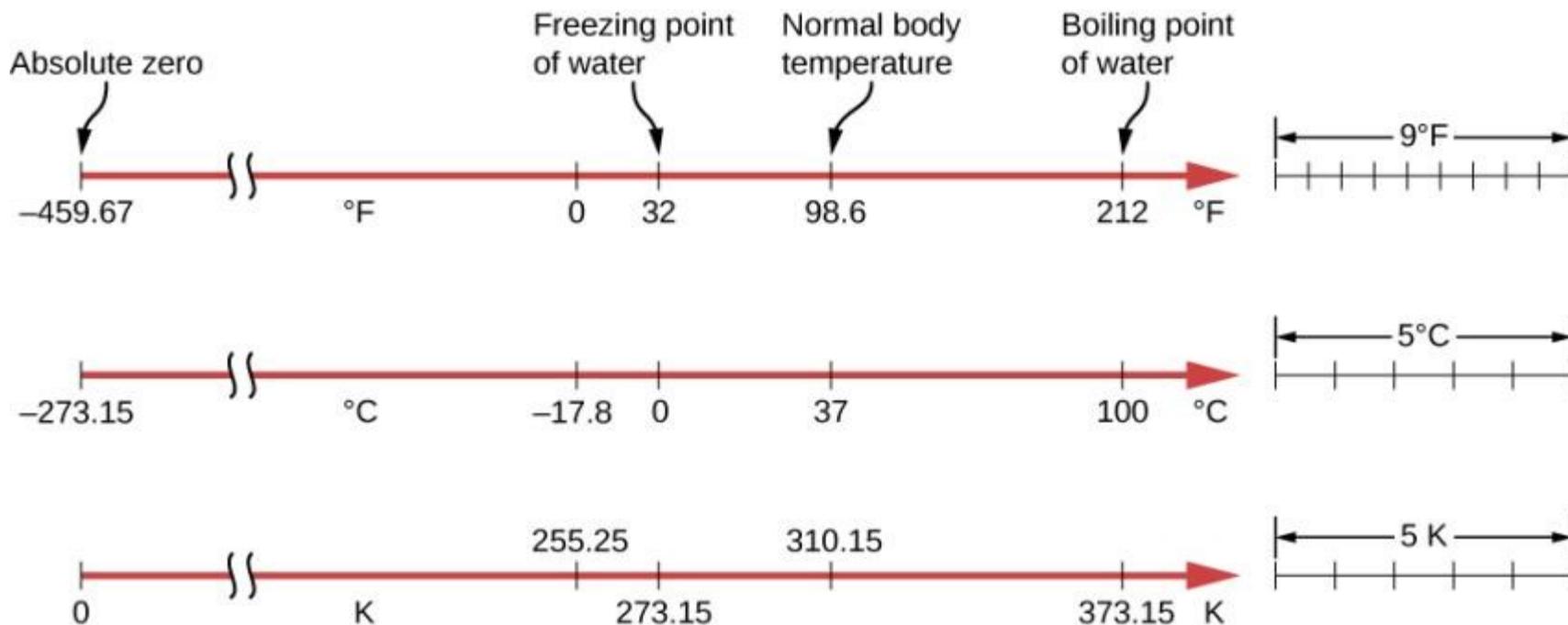
Introduction

FIGURE 1.2



If thermometer *A* is in thermal equilibrium with object *B*, and *B* is in thermal equilibrium with *C*, then *A* is in thermal equilibrium with *C*. Therefore, the reading on *A* stays the same when *A* is moved over to make contact with *C*.

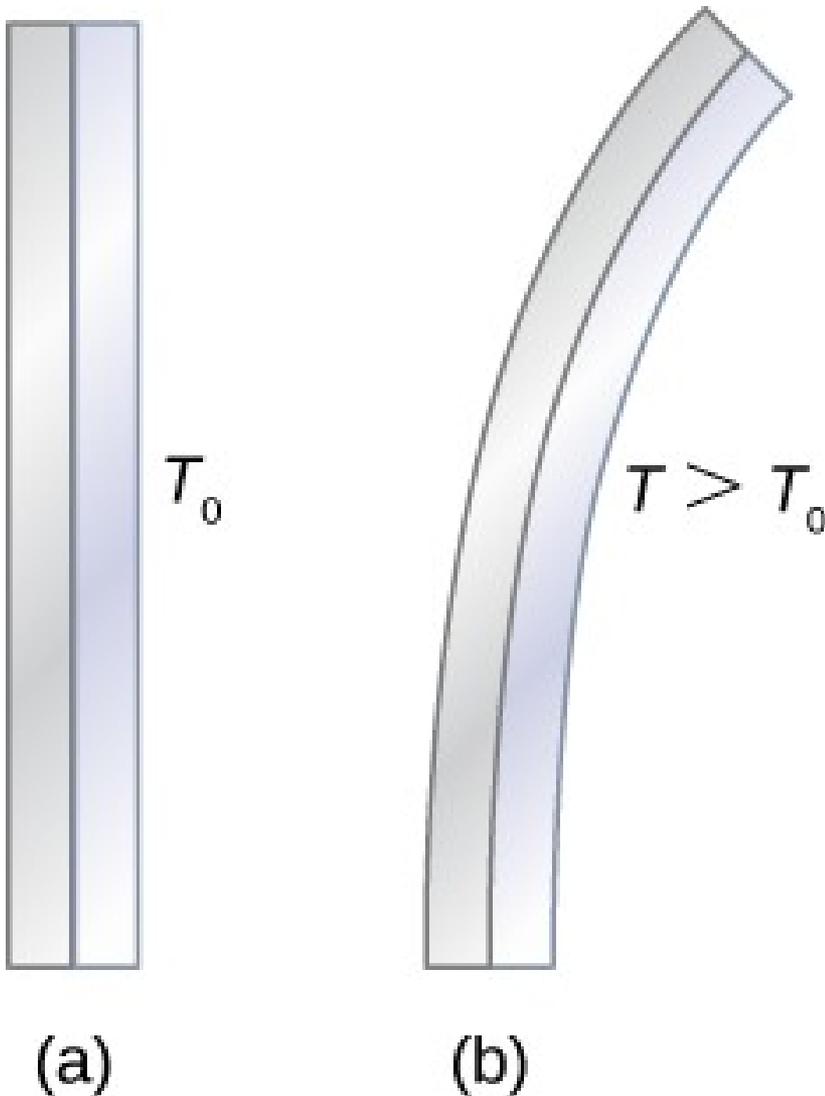
FIGURE 1.4



Relationships between the Fahrenheit, Celsius, and Kelvin temperature scales are shown. The relative sizes of the scales are also shown.

Thermal expansion

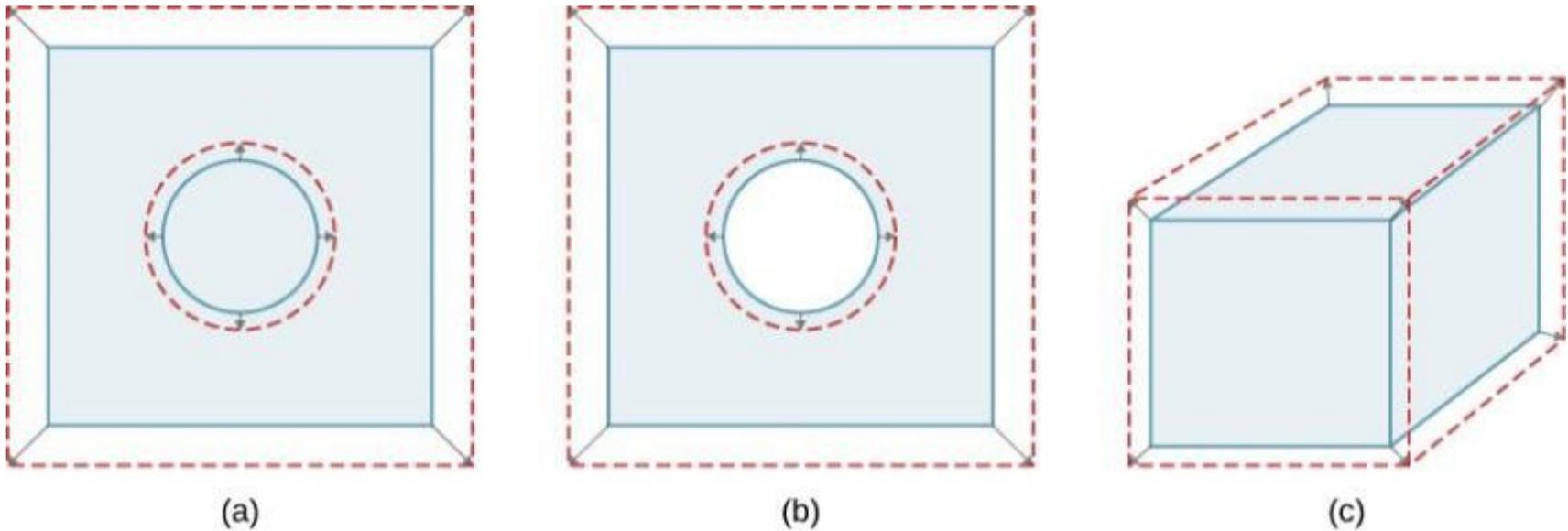
FIGURE 1.6



The curvature of a bimetallic strip depends on temperature.

- a) The strip is straight at the starting temperature, where its two components have the same length.
- b) At a higher temperature, this strip bends to the right, because the metal on the left has expanded more than the metal on the right. At a lower temperature, the strip would bend to the left.

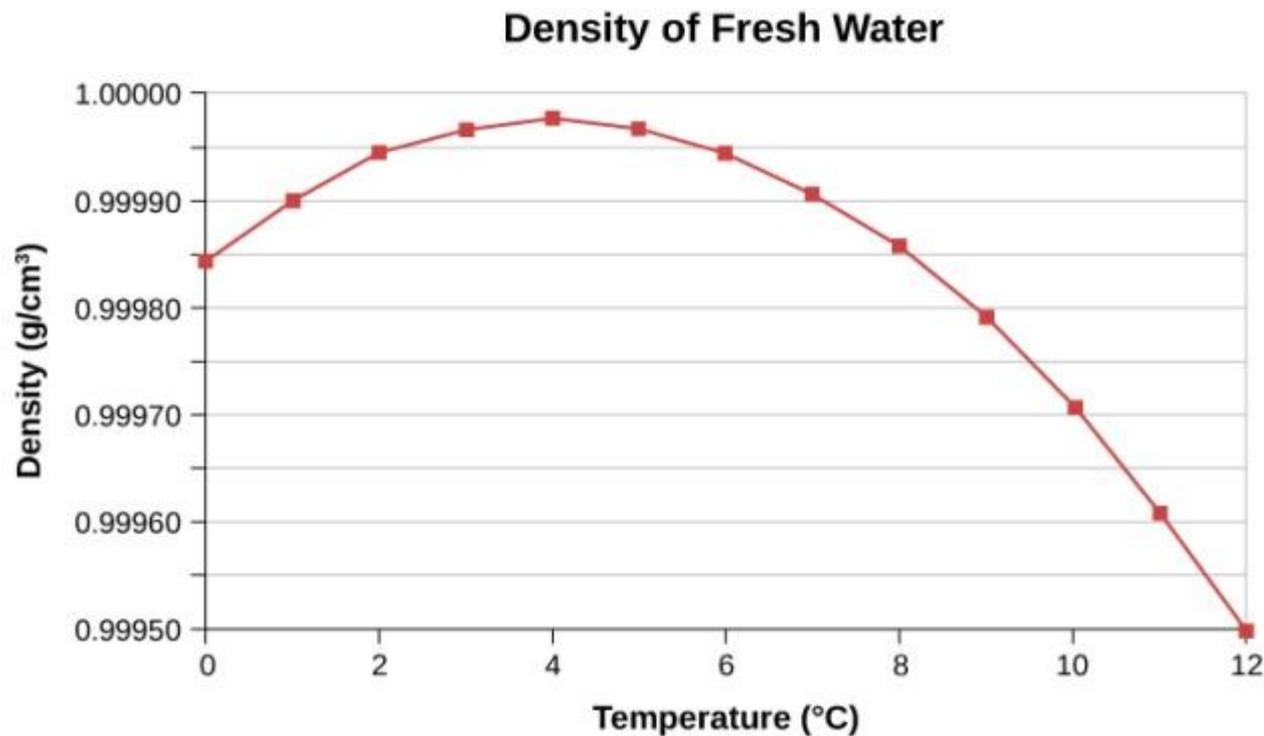
FIGURE 1.7



In general, objects expand in all directions as temperature increases. In these drawings, the original boundaries of the objects are shown with solid lines, and the expanded boundaries with dashed lines.

- a) Area increases because both length and width increase. The area of a circular plug also increases.
- b) If the plug is removed, the hole it leaves becomes larger with increasing temperature, just as if the expanding plug were still in place.
- c) Volume also increases, because all three dimensions increase.

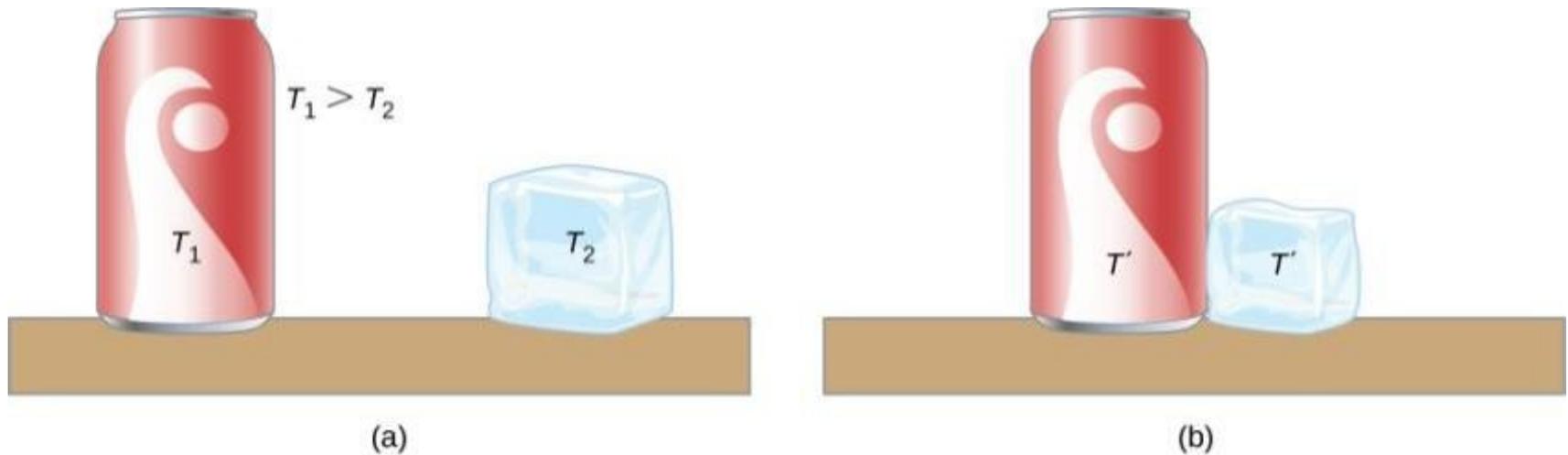
FIGURE 1.8



This curve shows the density of water as a function of temperature. Note that the thermal expansion at low temperatures is very small. The maximum density at 4 °C is only 0.0075% greater than the density at 2 °C, and 0.012% greater than that at 0 °C. The decrease of density below 4 °C occurs because the liquid water approaches the solid crystal form of ice, which contains more empty space than the liquid.

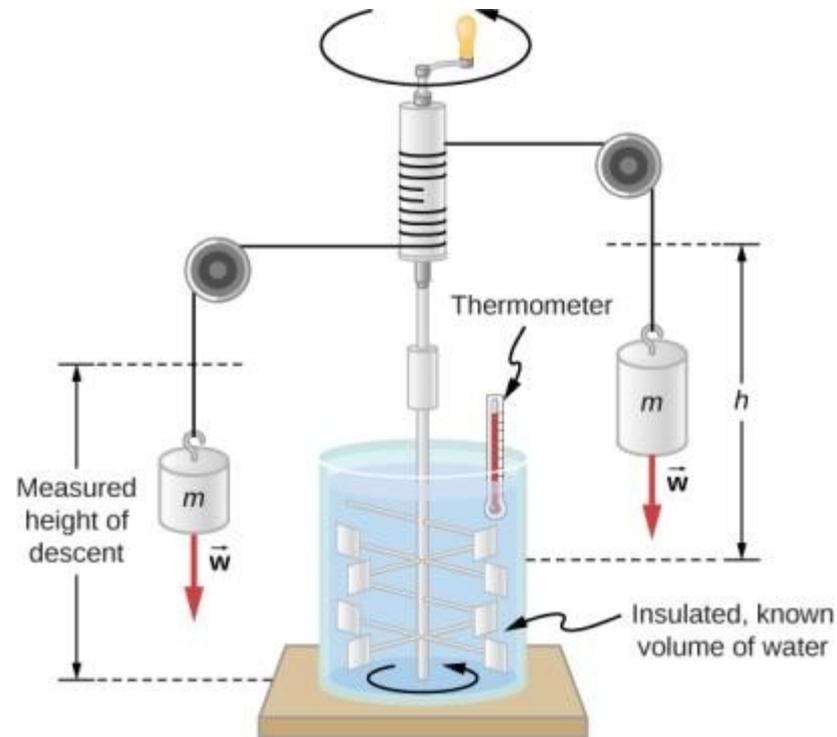
Equilibrium

FIGURE 1.9



- a) Here, the soft drink has a higher temperature than the ice, so they are not in thermal equilibrium.
- b) When the soft drink and ice are allowed to interact, heat is transferred from the drink to the ice due to the difference in temperatures until they reach the same temperature, T' , achieving equilibrium. In fact, since the soft drink and ice are both in contact with the surrounding air and the bench, the ultimate equilibrium temperature will be the same as that of the surroundings.

FIGURE 1.10



Joule's experiment established the equivalence of heat and work. As the masses descended, they caused the paddles to do work, $W = mgh$, on the water. The result was a temperature increase, T , measured by the thermometer. Joule found that T was proportional to W and thus determined the mechanical equivalent of heat.

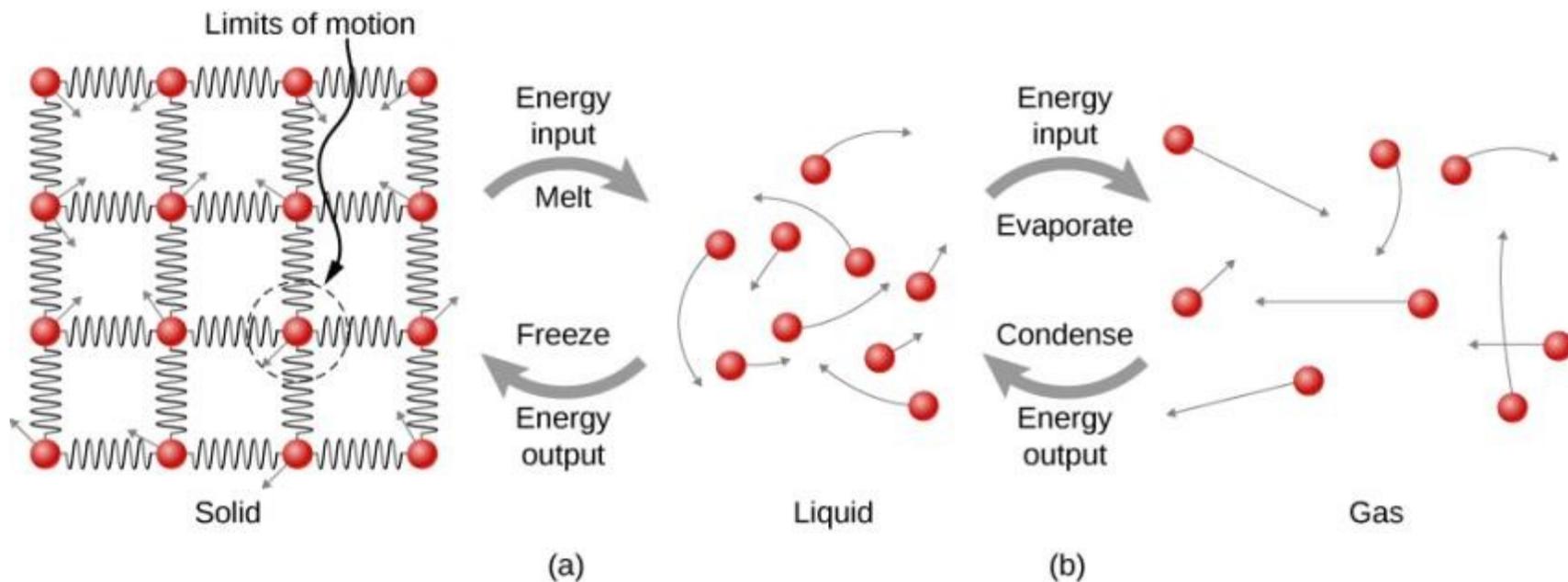
FIGURE 1.11



The smoking brakes on a braking truck are visible evidence of the mechanical equivalent of heat.

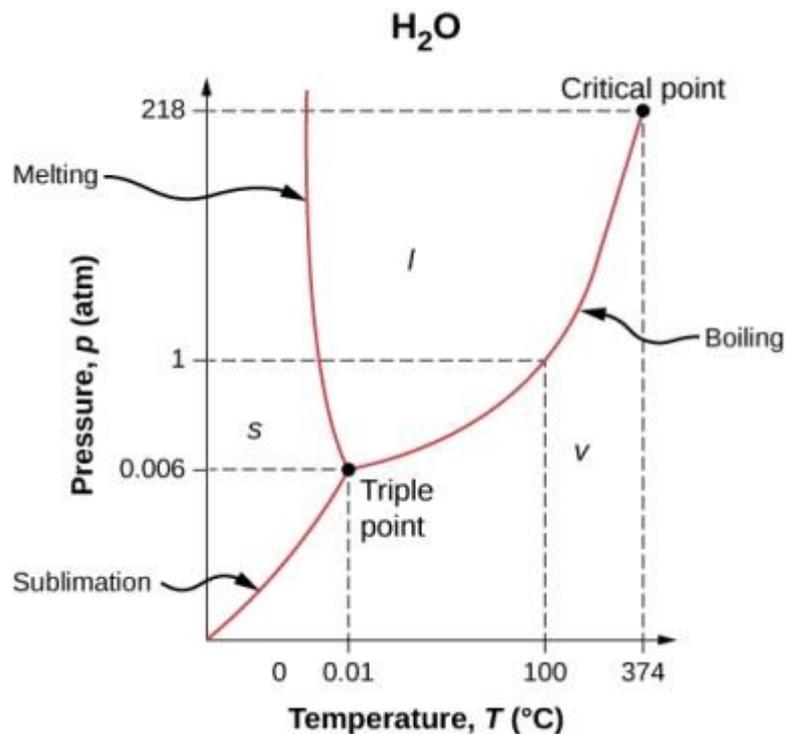
Phase transitions

FIGURE 1.18



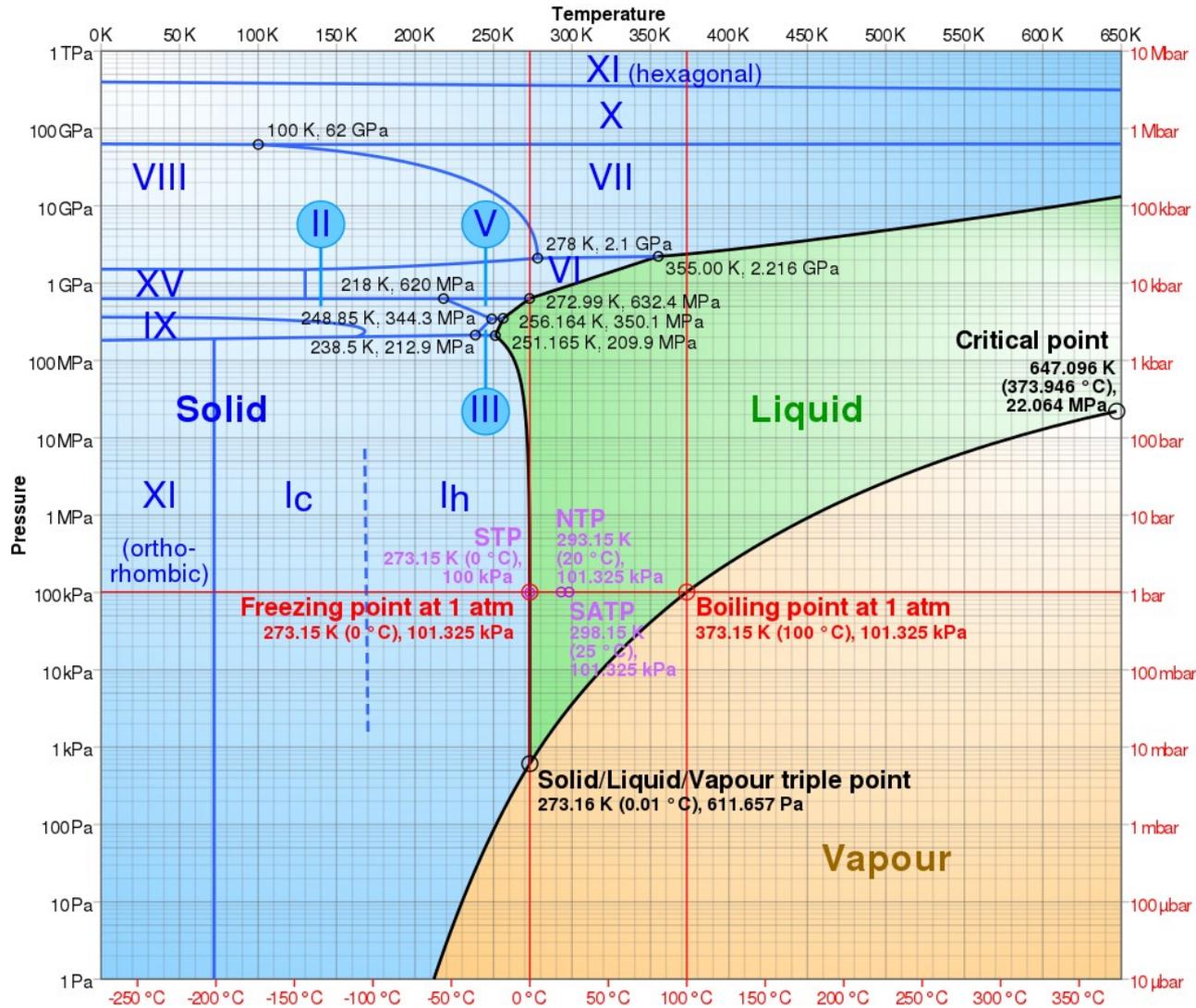
- Energy is required to partially overcome the attractive forces (modeled as springs) between molecules in a solid to form a liquid. That same energy must be removed from the liquid for freezing to take place.
- Molecules become separated by large distances when going from liquid to vapor, requiring significant energy to completely overcome molecular attraction. The same energy must be removed from the vapor for condensation to take place.

FIGURE 1.12



The phase diagram (pT graph) for water shows solid (s), liquid (l), and vapor (v) phases. At temperatures and pressure above those of the critical point, there is no distinction between liquid and vapor. Note that the axes are nonlinear and the graph is not to scale. This graph is simplified—it omits several exotic phases of ice at higher pressures. The phase diagram of water is unusual because the melting-point curve has a negative slope, showing that you can melt ice by *increasing* the pressure.

PHASES OF WATER



https://upload.wikimedia.org/wikipedia/commons/thumb/0/08/Phase_diagram_of_water.svg/220px-Phase_diagram_of_water.svg.png

FIGURE 1.13



(a)

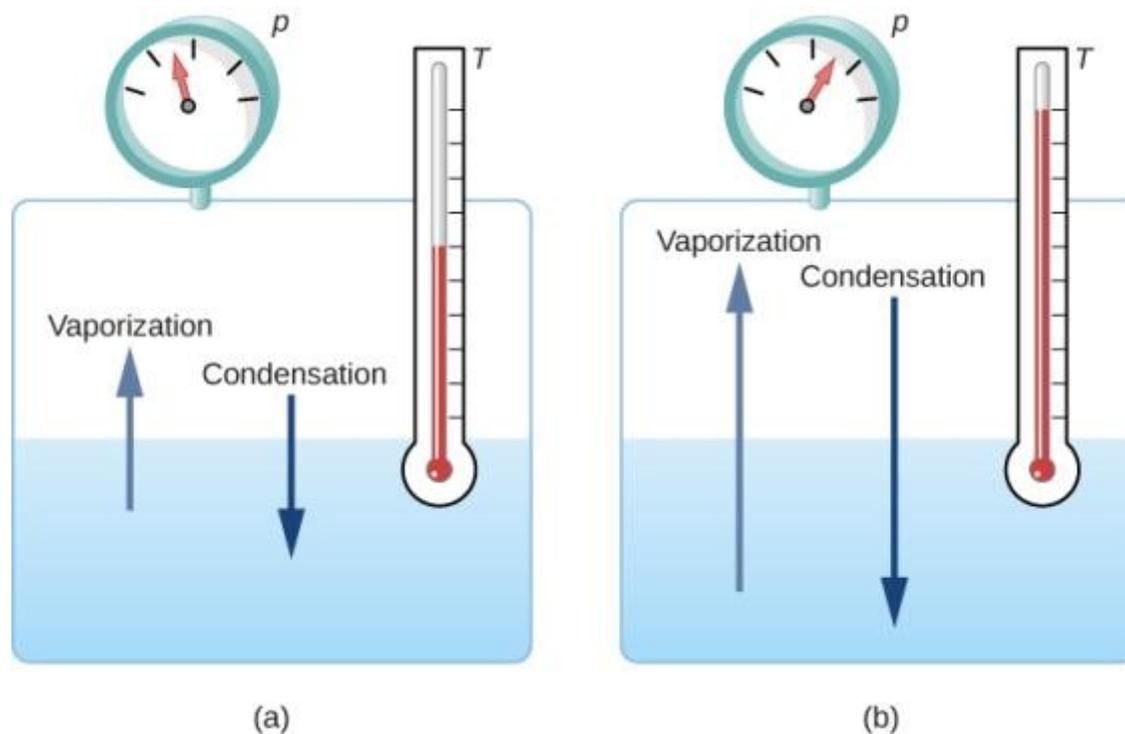


(b)

Direct transitions between solid and vapor are common, sometimes useful, and even beautiful.

- a) Dry ice sublimates directly to carbon dioxide gas. The visible “smoke” consists of water droplets that condensed in the air cooled by the dry ice.
- b) Frost forms patterns on a very cold window, an example of a solid formed directly from a vapor. (credit a: modification of work by Windell Oskay; credit b: modification of work by Liz West)

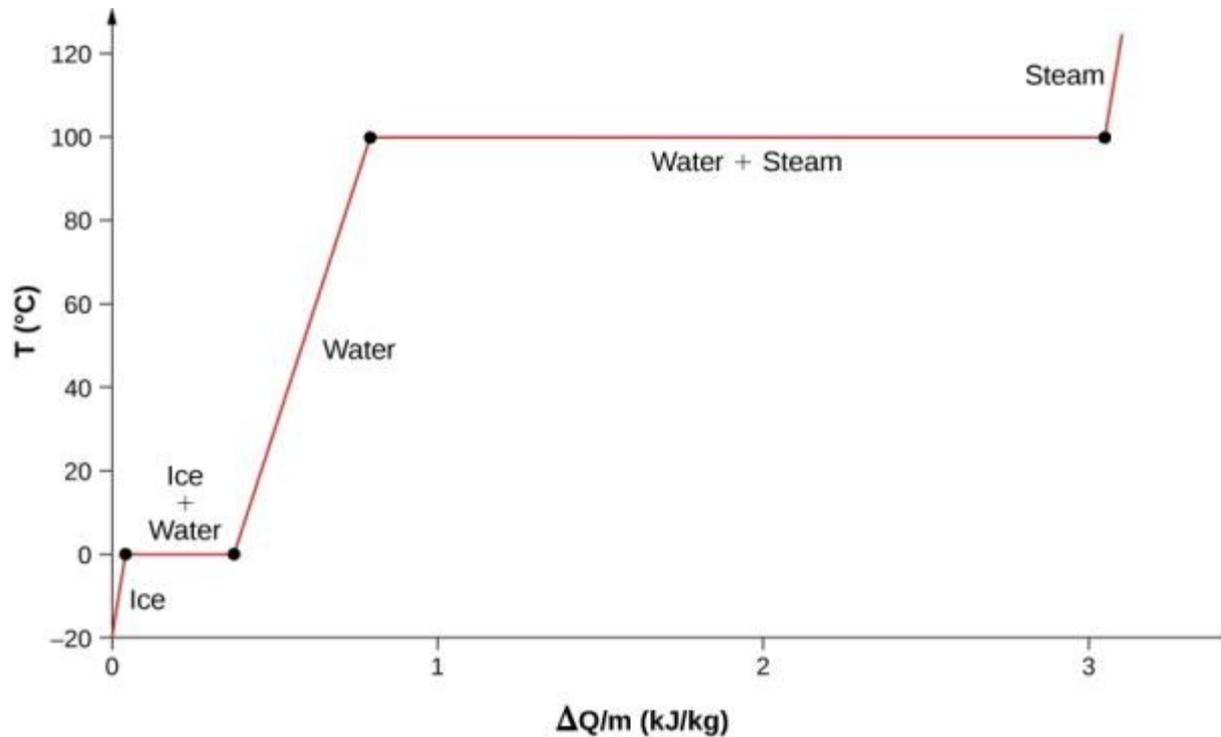
FIGURE 1.14



Equilibrium between liquid and gas at two different boiling points inside a closed container.

- The rates of boiling and condensation are equal at this combination of temperature and pressure, so the liquid and gas phases are in equilibrium.
- At a higher temperature, the boiling rate is faster, that is, the rate at which molecules leave the liquid and enter the gas is faster. This increases the number of molecules in the gas, which increases the gas pressure, which in turn increases the rate at which gas molecules condense and enter the liquid. The pressure stops increasing when it reaches the point where the boiling rate and the condensation rate are equal. The gas and liquid are in equilibrium again at this higher temperature and pressure.

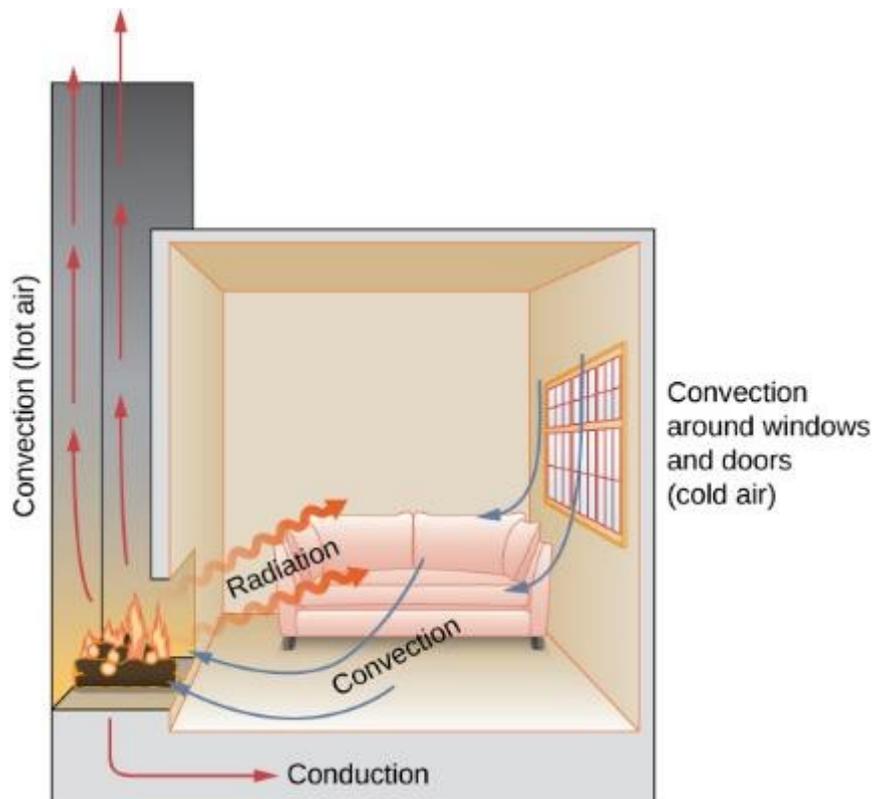
FIGURE 1.15



Temperature versus heat. The system is constructed so that no vapor evaporates while ice warms to become liquid water, and so that, when vaporization occurs, the vapor remains in the system. The long stretches of constant temperatures at 0 °C and 100 °C reflect the large amounts of heat needed to cause melting and vaporization, respectively.

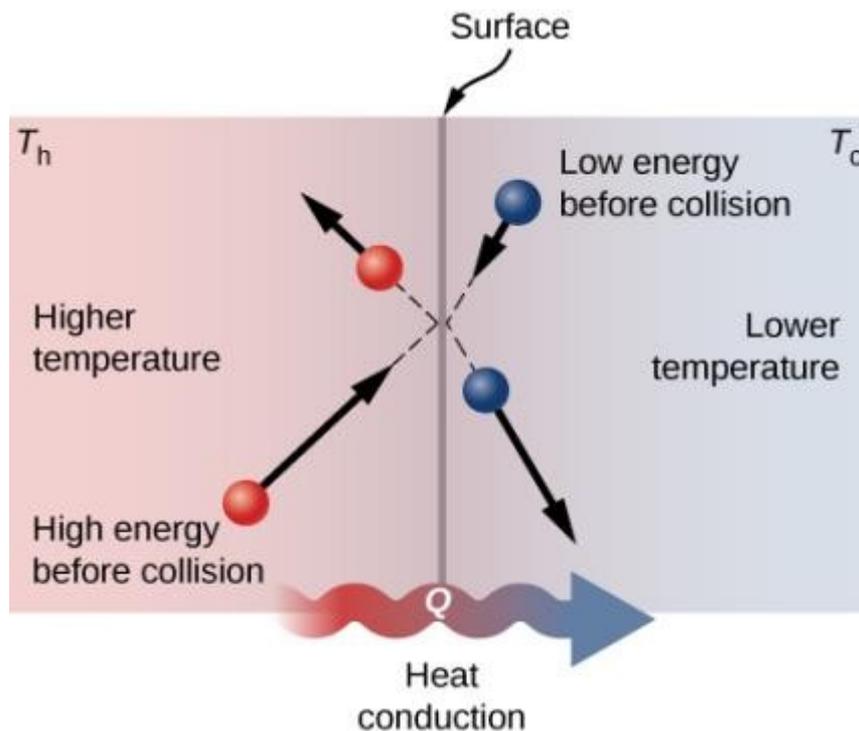
Heat transfer

FIGURE 1.19



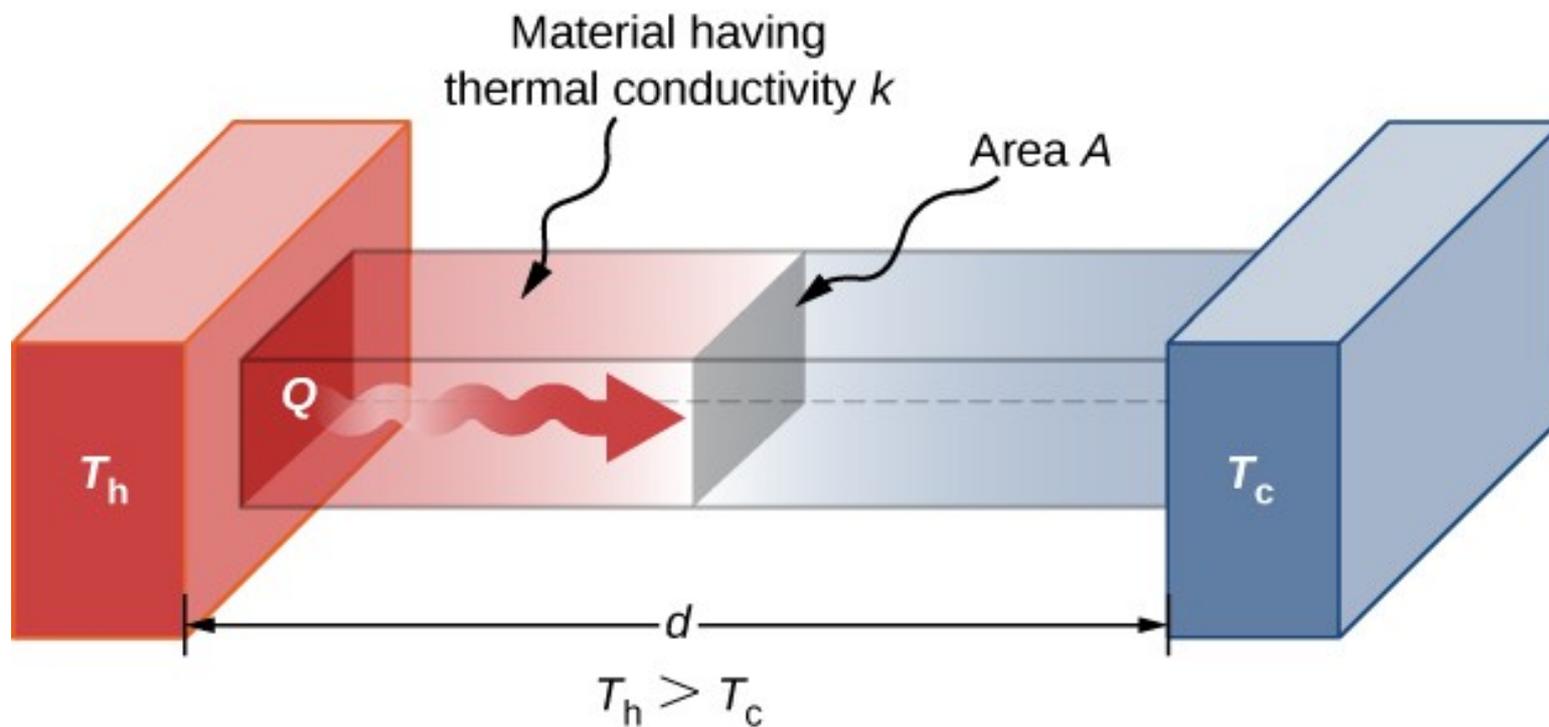
In a fireplace, heat transfer occurs by all three methods: conduction, convection, and radiation. Radiation is responsible for most of the heat transferred into the room. Heat transfer also occurs through conduction into the room, but much slower. Heat transfer by convection also occurs through cold air entering the room around windows and hot air leaving the room by rising up the chimney.

FIGURE 1.21



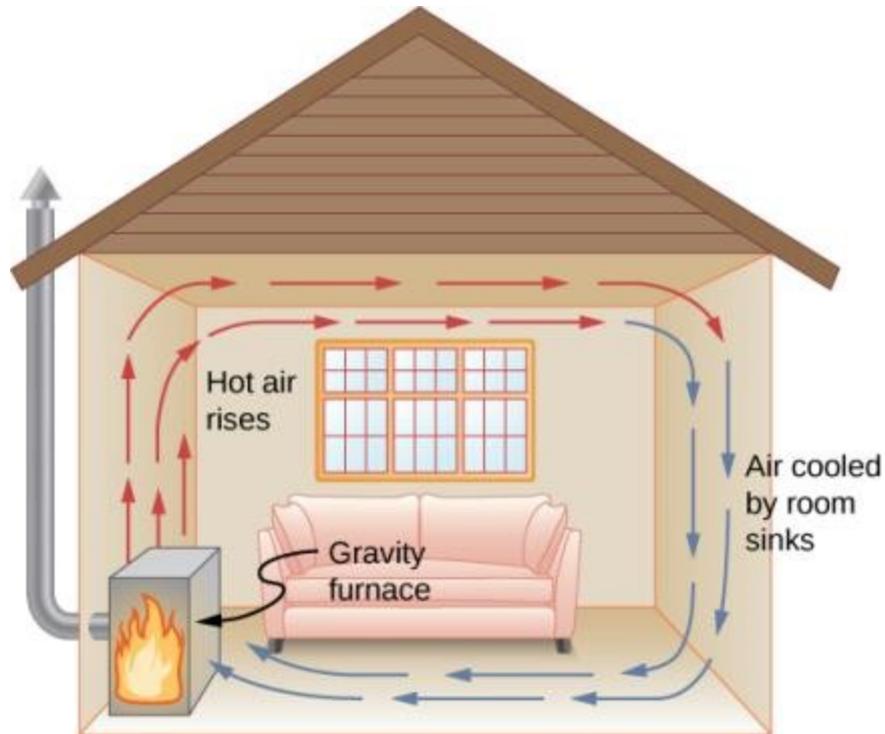
Molecules in two bodies at different temperatures have different average kinetic energies. Collisions occurring at the contact surface tend to transfer energy from high-temperature regions to low-temperature regions. In this illustration, a molecule in the lower-temperature region (right side) has low energy before collision, but its energy increases after colliding with a high-energy molecule at the contact surface. In contrast, a molecule in the higher-temperature region (left side) has high energy before collision, but its energy decreases after colliding with a low-energy molecule at the contact surface.

FIGURE 1.22



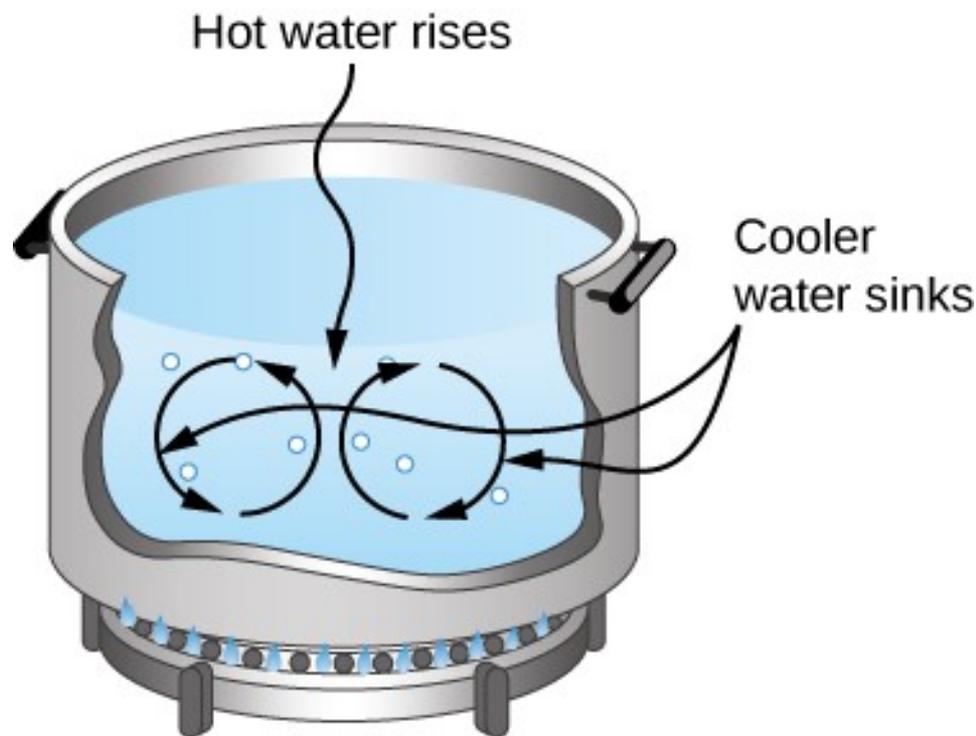
Heat conduction occurs through any material, represented here by a rectangular bar, whether window glass or walrus blubber.

FIGURE 1.24



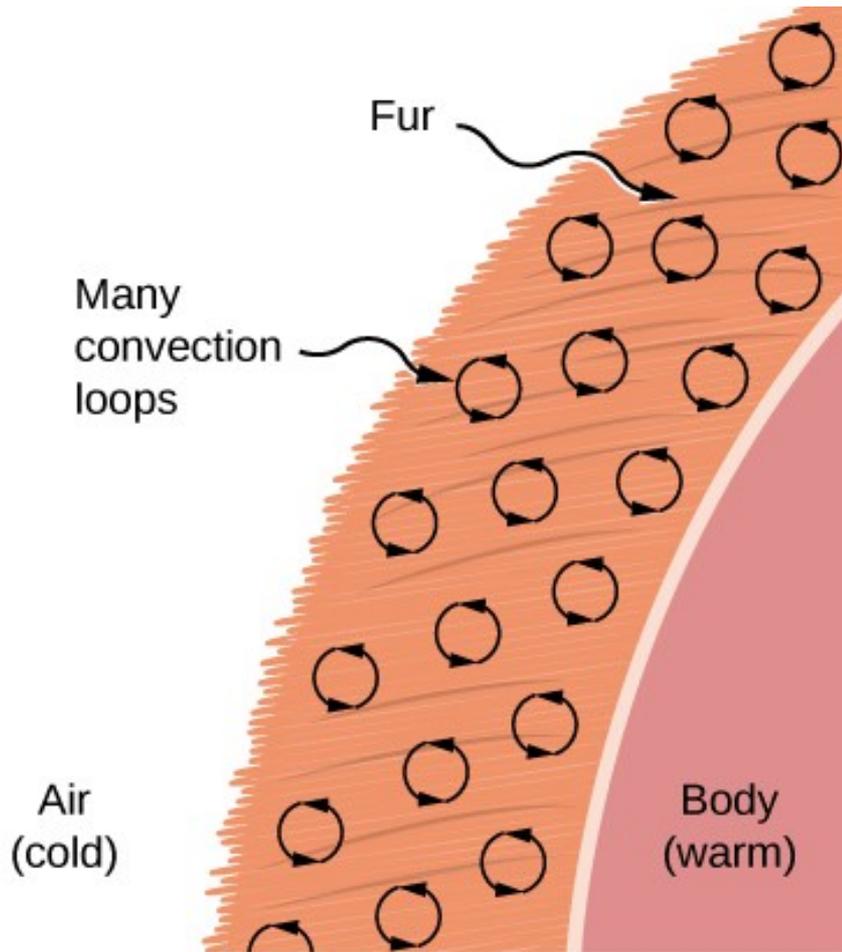
Air heated by a so-called gravity furnace expands and rises, forming a convective loop that transfers energy to other parts of the room. As the air is cooled at the ceiling and outside walls, it contracts, eventually becoming denser than room air and sinking to the floor. A properly designed heating system using natural convection, like this one, can heat a home quite efficiently.

FIGURE 1.25



Natural convection plays an important role in heat transfer inside this pot of water. Once conducted to the inside, heat transfer to other parts of the pot is mostly by convection. The hotter water expands, decreases in density, and rises to transfer heat to other regions of the water, while colder water sinks to the bottom. This process keeps repeating.

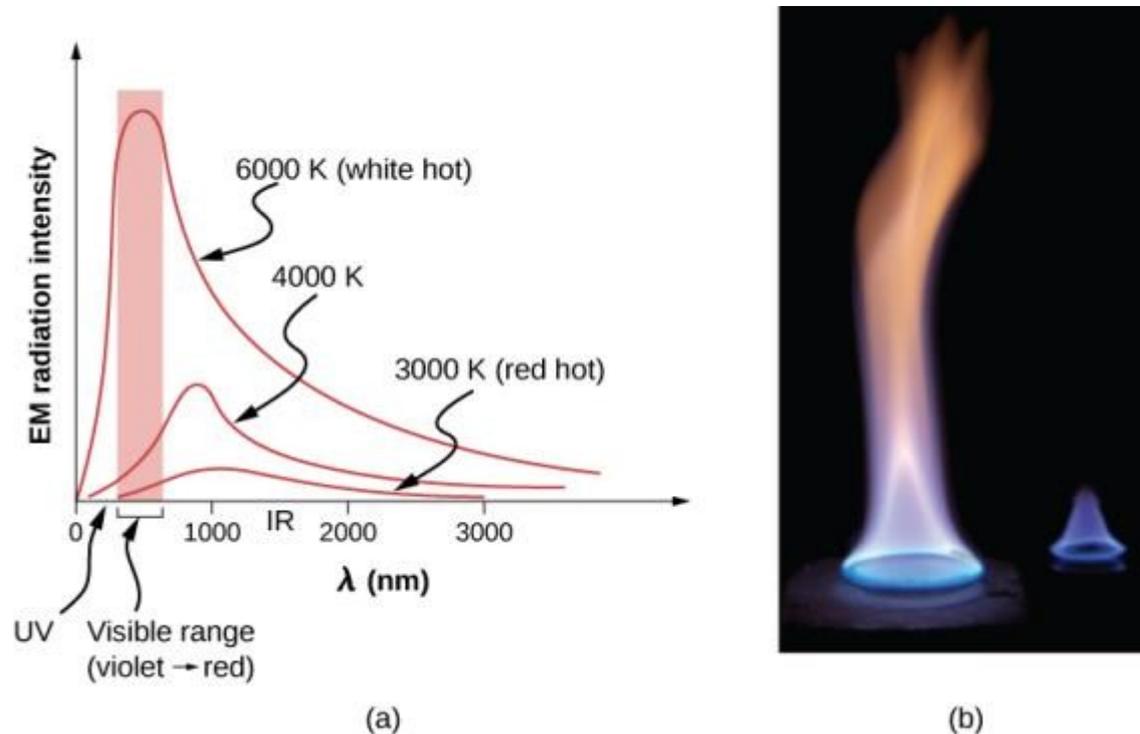
FIGURE 1.26



Fur is filled with air, breaking it up into many small pockets. Convection is very slow here, because the loops are so small. The low conductivity of air makes fur a very good lightweight insulator.

Radiation

FIGURE 1.29



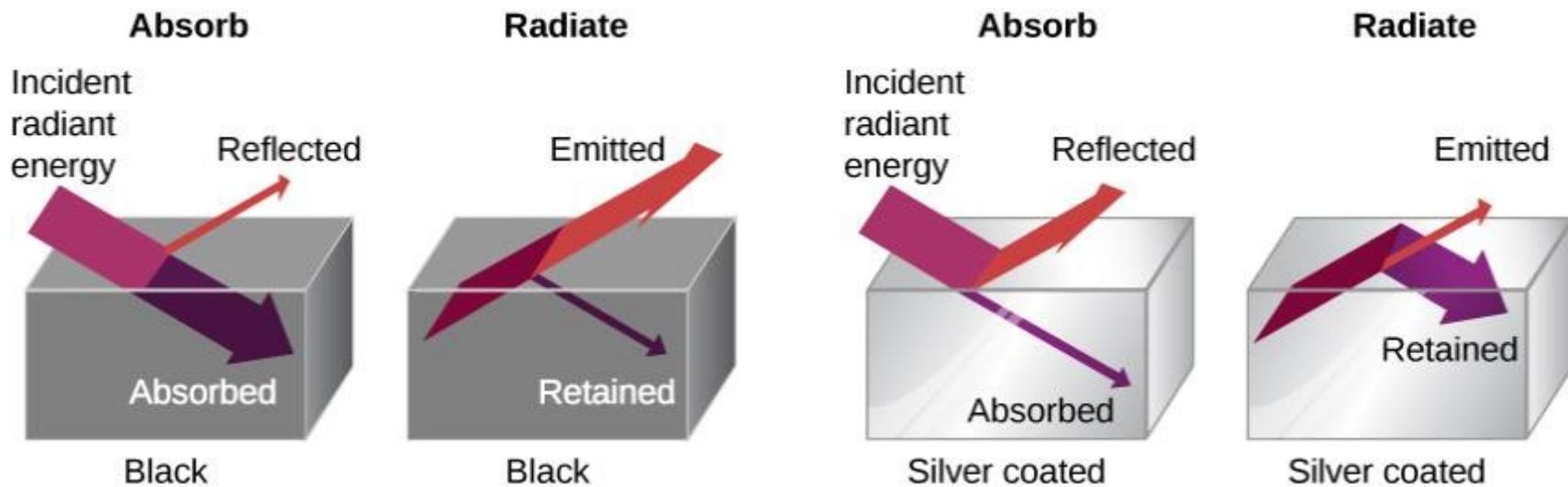
- a) A graph of the spectrum of electromagnetic waves emitted from an ideal radiator at three different temperatures. The intensity or rate of radiation emission increases dramatically with temperature, and the spectrum shifts down in wavelength toward the visible and ultraviolet parts of the spectrum. The shaded portion denotes the visible part of the spectrum. It is apparent that the shift toward the ultraviolet with temperature makes the visible appearance shift from red to white to blue as temperature increases.
- b) Note the variations in color corresponding to variations in flame temperature.

FIGURE 1.30



The darker pavement is hotter than the lighter pavement (much more of the ice on the right has melted), although both have been in the sunlight for the same time. The thermal conductivities of the pavements are the same.

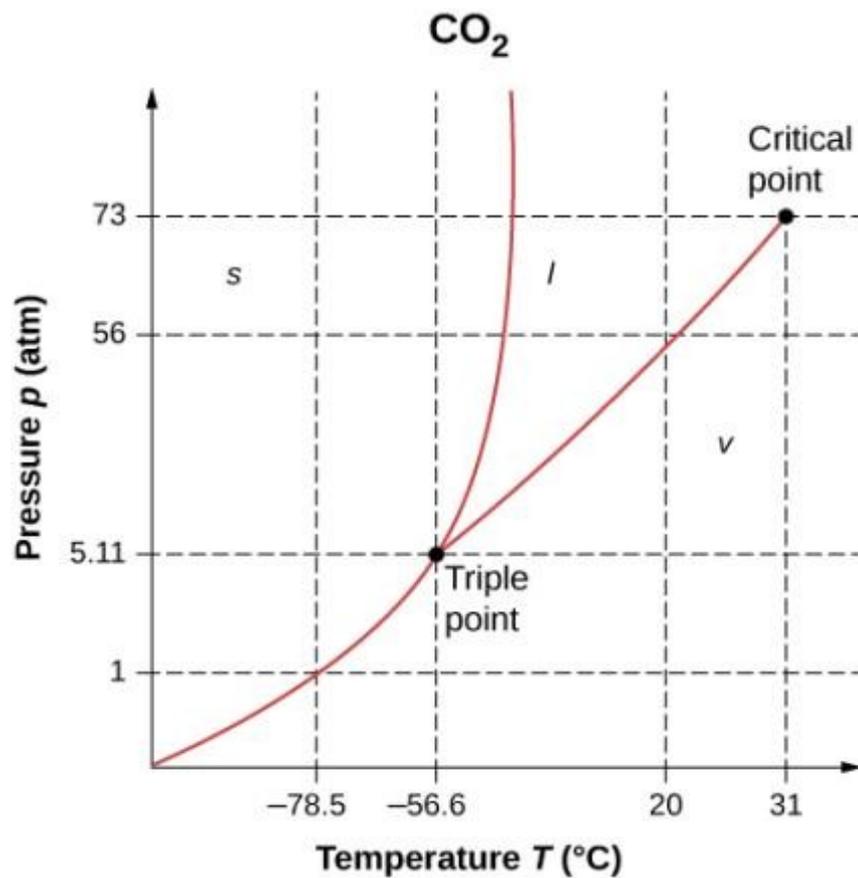
FIGURE 1.31



A black object is a good absorber and a good radiator, whereas a white, clear, or silver object is a poor absorber and a poor radiator.

Examples

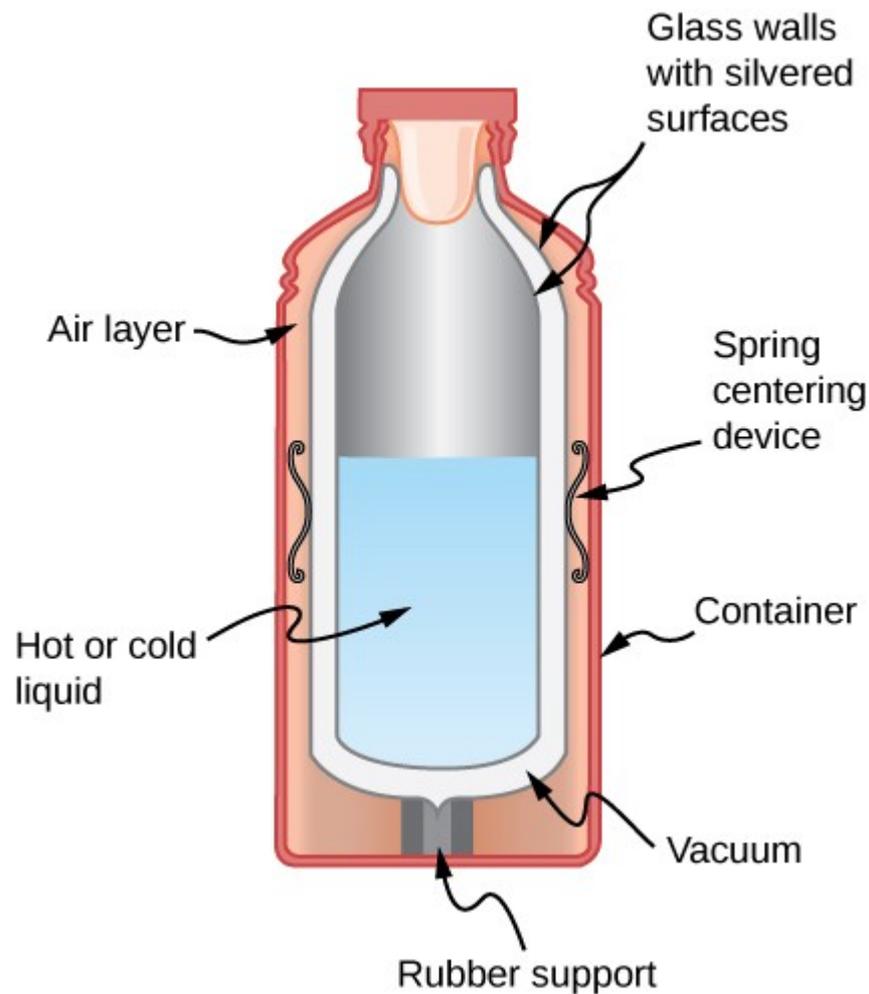
EXERCISE 16



As shown below, which is the phase diagram for carbon dioxide, what is the vapor pressure of solid carbon dioxide (dry ice) at (Note that the axes in the figure are nonlinear and the graph is not to scale.)

EXERCISE 30

Shown below is a cut-away drawing of a thermos bottle (also known as a Dewar flask), which is a device designed specifically to slow down all forms of heat transfer. Explain the functions of the various parts, such as the vacuum, the silvering of the walls, the thin-walled long glass neck, the rubber support, the air layer, and the stopper.





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